

Density Functional Theory Studies on Sulfur–Polyacrylonitrile as a Cathode Host Material for Lithium–Sulfur Batteries

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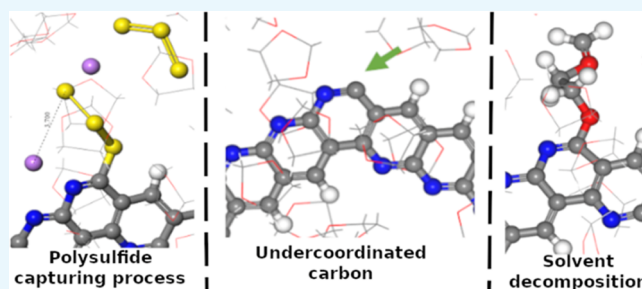


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ABSTRACT: Cyclized polyacrylonitrile, which can be obtained by vulcanization of polyacrylonitrile with sulfur, is an electron-conductive polymer that can be used as a host material in lithium–sulfur batteries. Using density functional theory, we investigated the interaction between a surrounding electrolyte and the polymeric sulfur–polyacrylonitrile (SPAN) electrode. In particular, we focused on different configurations, where the system contains 1,3-dioxane as a solvent and can have (i) polysulfide (PS) solvated in the electrolyte, (ii) a PS attached to the polymer backbone, (iii) lithium bis(trifluoromethanesulfonyl)imide (LiTFSI) as a salt dissolved in the electrolyte, and (iv) both PS and LiTFSI dissolved in the electrolyte. We found that the polymer, when having a hydrogen vacancy at a carbon atom (undercoordinated carbon) of the polymer backbone, is able to not only capture a PS from the electrolyte but also decompose and bind to the solvent and/or remove lithium from the PS. During this capturing process, the polysulfide might undergo S–S bond cleavage and recombination, accompanied by a charge transfer between the polysulfide and polymer. Thus, cyclized polyacrylonitrile not only is an interesting host material but also acts as an active material, together with sulfur, by capturing Li from the polysulfide.



INTRODUCTION

Lithium batteries are used for energy storage in different devices, such as cellphones and electric vehicles, and many other applications, both stationary and portable.^{1–6} However, the current specific capacity (150 Wh·kg⁻¹) of Li-ion batteries needs to be augmented to supply the increasing demands, for instance in electric vehicles.^{7,8} Here, lithium–sulfur batteries (LiSBs) are promising future candidates due to their high theoretical specific capacity of 2567 Wh·kg⁻¹, low toxicity, and low cost of the graphite/sulfur cathode.^{9–13} However, LiSBs have some challenges that still have to be overcome, such as blockage of the active material in the cathode, the formation of isolator layers for electron passage, volumetric expansion, and the so-called polysulfide (PS) shuttle.^{14–16} During cycling of LiSBs, long-chain polysulfides (Li₂S_x, 3 ≤ x ≤ 8) can dissolve into and shuttle through the electrolyte, facilitating degradation of the anode and self-discharge.^{17–20}

Since sulfur is a nonconductive material, graphite is usually used as a host material in LiSBs, providing the necessary electronic conductivity. Many strategies to reduce the degradation and fading cyclability of LiSBs, which is promoted by the migration of the PS to the anode, are based on changing the architecture of this graphite, for instance, by using carbon nanotubes, micropores, hollow carbon spheres, or by encapsulating sulfur inside a polymer as a yolk, where the polymer shell allows Li-ion transportation.^{21–37} Additives to the host material are also used to force the dissolved polysulfides to adsorb again into the electrolyte, such as

magnesium oxide particles added to the graphite.^{38–42} The graphite can also be modified by doping, e.g., with nitrogen, to chemisorb the PSs.^{43–45} Other attempts have tried to reduce the PS shuttle effects by using membranes that decrease the amount of PSs reaching the anode or insertion of an additive to the electrolyte to actively engineer the solid–electrolyte interphase (SEI) at the anode, finally protecting the system against PS degradation.⁴⁶ The SEI can also be produced at the cathode by the attack of some salts (e.g., LiBr) on the electrolyte, leading to an SEI at the cathode that reduces PS migration.^{47–50}

Alternative materials to graphite as host materials in LiSBs that also improve the cycling performance are polymer–sulfur composites, of which polyacrylonitrile (PAN) is an interesting candidate for the cathode side.^{51–56} Vulcanization of PAN with sulfur leads to cyclization of the PAN polymer, producing π -conjugated motifs as well as sulfur chains that are chemically bound to C atoms of the PAN polymer.^{57–59} During PAN synthesis, it is possible to observe a reduction of the hydrogen content and, at the same time, the formation of hydrogen

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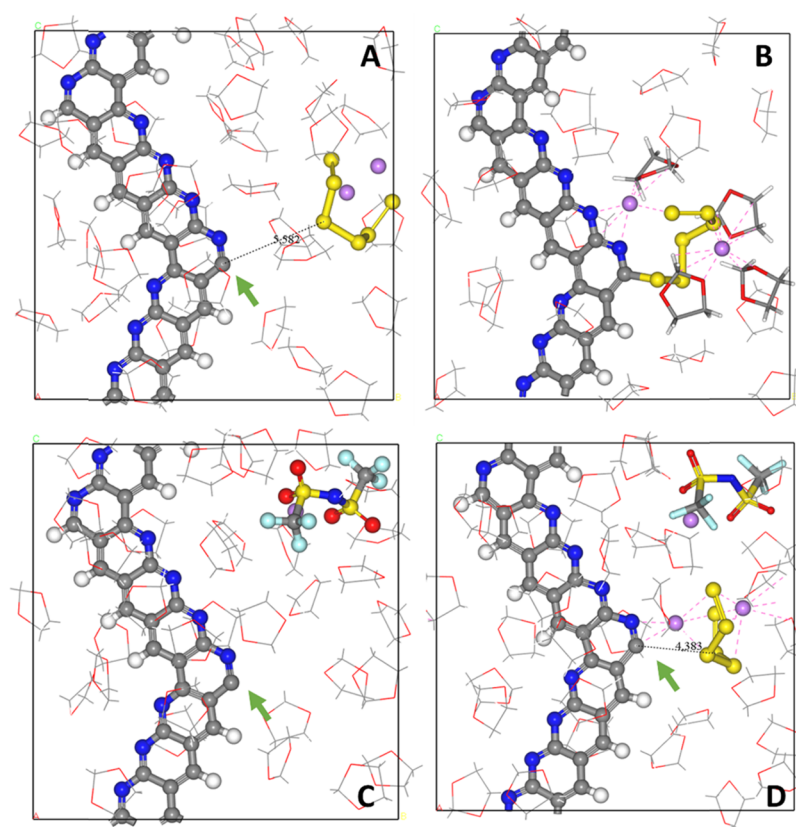


Figure 1. Initial structure models used to investigate the interactions between PS and LiTFSI with a single site of c-PAN. (A) Single C_{uc} surrounded by a PS-containing electrolyte. (B) Site in c-PAN which is considered initially occupied by PS. (C) Single C_{uc} surrounded by a LiTFSI-containing electrolyte. (D) Single C_{uc} surrounded by an electrolyte containing both LiTFSI and PS. Color coding: hydrogen (white), lithium (purple), carbon (gray), nitrogen (blue), oxygen (red), fluorine (cyan), and sulfur (yellow). Green arrows point to a carbon with a missing hydrogen, representing a C_{uc} .

sulfides and N–C–S units, indicating the formation of 2-pyridylthiolates and thioamides.^{57,60} After this synthesis, the polymeric structure contains π -conjugated units, enabling the usual identification as cyclized polyacrylonitrile (c-PAN). Regarding its use in batteries, the performance of this polymer as the host material is affected by the temperature of sulfur–polyacrylonitrile (SPAN) vulcanization, the sulfur content, and the solvent composition, indicating that the synthesis process strongly affects the behavior of the LiSB. Chen et al. suggested that dehydrogenation of PAN first creates the cyclized structure (c-PAN) and then produces uncoordinated carbons on the c-PAN backbone.⁶¹ Moreover, Archer et al.⁶² observed oscillations in the number of C–S bonds during cycling, which was interpreted as a rupture process between the polymer and the sulfur chain during lithiation.

In this work, we have used density functional theory (DFT) to resolve the structure of sulfur–polyacrylonitrile^{58,59} and to investigate the PS shuttle effect in the anode as well as graphite cathodes.^{63–70} We further considered the situation that sulfur binds to an undercoordinated carbon such as in c-PAN. Here, we used ab initio molecular dynamics (AIMD) to study the interaction of the host material with a polysulfide-containing electrolyte. AIMD is commonly applied for investigation on LiSB (using a time scale of picoseconds), due to its high accuracy, both to describe electron distribution and to supply energy to overpass the active barrier in a reaction. In contrast, classical methods and static calculations are not able to provide these outcomes at the same time.^{71–74} AIMD can give insights

into the behavior of the cathode, and the structure of adsorbed PSs and help answering the question of whether c-PAN is capable of capturing PSs from the electrolyte. As c-PAN as a host material shows a higher cyclability than graphite, one of our aims was to understand the possibility of chemisorption and incorporation of the PSs into the c-PAN network, as well as the role and influence of the electrolyte on this process.

COMPUTATIONAL METHODS

DFT calculations were performed with the Vienna Ab initio Simulation Package (VASP)^{75,76} employing plane-wave basis sets^{77,78} with an energy cutoff of 400 eV, while electron–core interactions were described by projector-augmented wave (PAW) pseudopotentials.^{79,80} For integrations in the reciprocal space, the Brillouin zone was sampled with a Monkhorst–Pack mesh⁸¹ of $2 \times 2 \times 2$ k -points. The electronic structure was relaxed until the energy was converged up to an energy change of 10^{-4} eV. Geometries were optimized using the conjugate-gradient (CG) approach with a force convergence criterion of 10^{-3} eV/Å. Gaussian smearing with a width of 0.05 eV was utilized. Ab initio molecular dynamics (AIMD) simulations were carried out in the canonical ensemble (NVT) at 330 K using time steps of 1 fs, while the Nosé thermostat was used to control temperature oscillations during the simulation with a Nosé mass parameter^{82–84} of 0.5. Charge analyses were performed on the basis of Bader charges⁸⁵ obtained with a denser reciprocal grid of $4 \times 4 \times 4$ k -points. Finally, in all studies, van der Waals (vdW) dispersion corrections were

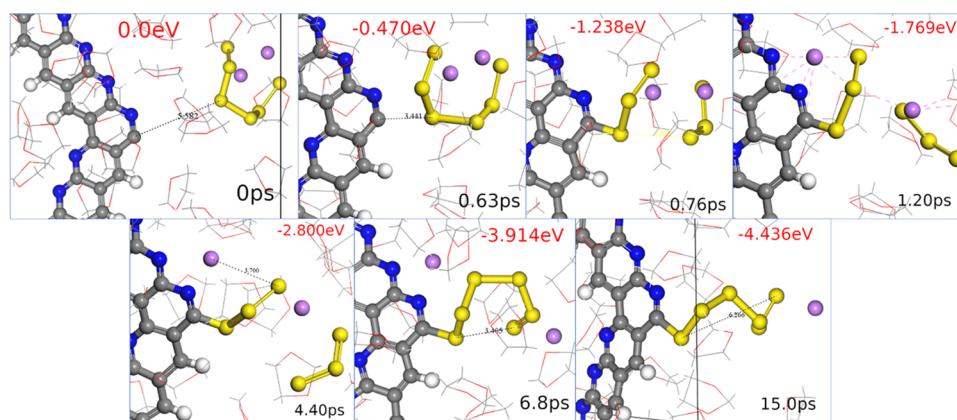


Figure 2. Screenshots along with the AIMD simulation showing the capturing of PS dissolved in the electrolyte by the c-PAN. The color code is as described in Figure 1.

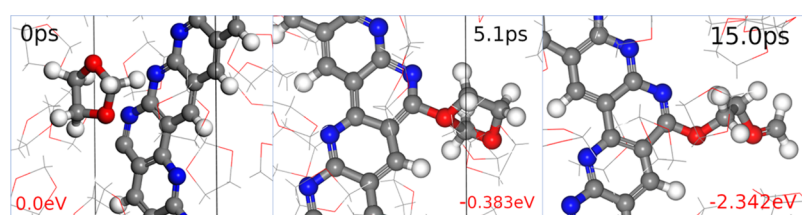


Figure 3. Decomposition mechanism of the solvent in the presence of a C_{uc} of c-PAN. The color code is as described in Figure 1.

included by the DFT-D3 approximation in the Becke–Johnson formulation.⁸⁶ To estimate individual bond strengths, the Crystal Orbital Hamilton Population (COHP) and Crystal Orbital Overlap Population (COOP) were calculated using the LOBSTER program.^{87,88} VESTA⁸⁹ software was used for visualization of the charge isosurface and visual molecular dynamics (VMD)⁹⁰ was used for visualization of the molecular structures.

To build up the different systems, the electrolyte molecules were preoptimized individually using the Gaussian09 (G09) package⁹¹ with the B3PW91 hybrid functional and a 6-311++G(p,d) basis set,^{92,93} allowing a better initial configuration to run in periodic boundary conditions when using VASP. Afterward, the full system (i.e., c-PAN with electrolyte) was built such that the individual electrolyte molecules were first kept rigid in their preoptimized structures, while c-PAN was optimized based on the consistent-valence forcefield (CVFF).⁹⁴ Finally, this provided the initial configurations for the subsequent AIMD simulations. This procedure allows saving computational resources and eliminating drifts in both energy and temperature at the beginning of the simulation.

As a solvent, we used 1,3-dioxolane (DOL) that contained one molecule of PS (Li_2S_6), for which the electrolyte density was set to be 1.06 g/cm³. The literature indicates that in the PAN structure, sulfur chains of five and eight atoms (S_n , $5 \leq n \leq 8$) are stable.^{58,59} Therefore, as a compromise for the present studies, we used polysulfides containing six sulfur atoms in the chain. All systems were studied in 15 Å × 18 Å × 18.5 Å simulation boxes, where the polymer backbone was oriented along the *z*-direction.

Figure 1 summarizes the different initial configurations for our AIMD studies that were used to investigate the interaction between PS, the electrolyte, and c-PAN. Here we successively increased the complexity of the system. In the first two cases (Figure 1A,B), the interaction of the PS (Li_2S_6) with c-PAN

was investigated by first considering the detached case (Figure 1A) and afterward binding the PS to one of the undercoordinated carbon atoms of the backbone (hereafter, this undercoordinated site is called C_{uc} ; see Figure 1A). The C_{uc} position was selected based on where the hydrogen in c-PAN is the most thermodynamically favorable to be removed. With the third and fourth systems, the role of the added salt lithium bis(trifluoromethanesulfonyl)imide (LiTFSI; Figure 1C) and finally the combined system where both PS and LiTFSI are in the simulation box were considered (see Figure 1D). The last two cases contain an undercoordinated carbon site, C_{uc} .

To highlight the different processes or reactions observed during the simulations, individual snapshots along the AIMD were selected. To further analyze the energetics between c-PAN and PS, for each selected frame an additional subsequent energy minimization was performed in such a way that the c-PAN and PS structures were kept fixed, while only the electrolyte molecules were geometrically optimized (see Figures 2 and 3, as well as Figures S1 and S2). This procedure provided direct insights into the energetics during the reaction processes between PS and c-PAN but reduced temperature-induced fluctuation effects in the solvent.

RESULTS AND DISCUSSION

To study the interaction of PS and c-PAN, our simulation model systems were constructed such that a single hydrogen termination on the carbon backbone of the c-PAN host material was removed (see Figure 1). This is motivated by the experimental observations that vulcanization of PAN leads to the formation of conjugated π -units, the production of H_2S , the formation of C–S bonds, and the creation of uncoordinated carbon atoms at the c-PAN backbone.^{57–59,61,95} Moreover, since some salts can remove hydrogens from the solvent surrounding the cathode,^{47–50} there is also the possibility that those salts remove hydrogens from the SPAN

backbone during cycling, producing an undercoordinated carbon on c-PAN. During cycling, lithium will react with the sulfur present in the SPAN and form a PS soluble in the electrolyte; thus, PS might also be present in the electrolyte. Therefore, we also investigated the interaction between c-PAN and an electrolyte molecule. Overall, the following different situations were considered: (1) an initially ringlike PS interacts with a C_{uc} (i.e., missing H-saturation) site of c-PAN; (2) this particular site of c-PAN remains unoccupied, i.e., there is a missing H atom in the backbone (C_{uc}) and the electrolyte contains only DOL and LiTFSI; (3) the C_{uc} site is in contact with the electrolyte that contains only DOL and a PS molecule being close to (5.5 Å) this C_{uc} (thus has a strong tendency to interact); and finally, (4) the electrolyte contains the solvent, PS, LiTFSI, and a C_{uc} on the c-PAN polymer. Additionally, we also positioned the LiTFSI closer to the C_{uc} site (see Figure S3) to investigate possible reactions between the polymer backbone and the salt and at the same time reducing the possibility of C_{uc} to react with the solvent.

In the model shown in Figure 1B, the PS has initially a ringlike structure connected to the c-PAN. During the simulation, the c-PAN receives a Li atom that was previously solvated by the PS. As observed in Figure S1, the Li atom migrates through the c-PAN, increasing its distance to the PS (from 2.8 to 5.4 Å). In the direct vicinity of the Li atom removed from the PS, there are N atoms from the c-PAN as well as O atoms from different solvent molecules that form the remaining solvation shell. Consequently, this Li has four nearest-neighbor atoms originating from different solvent molecules and c-PAN. Once Li is removed from the PS, the PS molecule elongates, i.e., the distance between S in the edge of the PS and the S atoms that form the connection to the polymer changes from 3.8 to 5.7 Å (comparing initial and final structures). Also, the other Li atom remaining at the PS moves to the edge of the PS chain (not connected to the polymer), and the first solvation shell of this Li atom is surrounded by one S atom and O atoms from the surrounding solvent molecules. This indicates the capacity of the polymer to interact with Li not only as a host material for sulfur but also by consuming Li, consequently acting as an active material during battery cycling. This behavior was also suggested by Wang et al.⁹⁶ who argued on the absorption of Li ions by the polymer backbone.

When the PS is initially dissolved into the electrolyte, the C_{uc} site of c-PAN is able to catch the PS molecule from the electrolyte (see Figure 2). During the first step of this capturing process, the PS remains intact with its ringlike structure, while the distance between the PS and the C_{uc} reduces from 5.8 to 3.4 Å. At a certain distance, the S atom closest to the C_{uc} has one of its S–S bonds broken; in this case, we observe a fragmentation that creates two polysulfides with three S atoms on each chain. While one of the PS chains is attached to the polymer, the other is kept by the Li atom shared between the two chains. With time, one of the Li atoms moves to the N atoms of the polymer, interacting electrostatically with one PS fragment. The other Li stays between both PS fragments, of which one is connected to the C_{uc} , while the other dissolves into the electrolyte. Subsequently, the bound PS detaches, resulting in Li being solely solvated by N atoms of the c-PAN and oxygen atoms from solvent molecules. This allows the attached PS fragment to reconnect to the other PS chain, forming a ringlike structure attached to the system. Finally, the PS elongates, assuming the same structure as in the

system with the initially attached PS (Figure 1B). The total energy change during the capturing process by a C_{uc} site is around -4.4 eV, which can be divided into different steps. The capturing and fragmentation of the PS into the C_{uc} is around -1.2 eV, producing two S_3 fragments. The removal of Li ions from the PS to c-PAN is around -1.6 eV, the fragments recombination is around -1.1 eV, and the PS reconfiguration is around -0.5 eV. Several intermediary steps are shown in Figure 2. C_{uc} can capture PS from the electrolyte, but the process may also create smaller PS fragments, whose sulfur chains recombine to produce longer chains. The simulations indicate that C_{uc} tends to capture the closest S atom from the PS, which results in fragmentation and recombination of the PS. Rotation of the PS was not observed setting the PS at 5.8 Å, but the PS should also be captured by the chain edge without producing fragmentation and recombination.

When both LiTFSI and PS are present in the electrolyte and not connected to the polymer, the PS is also attracted by the C_{uc} site, leading now to a fragmentation that produces S_2 as well as S_4 (Figure S2). Without the presence of c-PAN, the fragmentation of the PS is not thermodynamically favorable⁶³ but becomes favorable in the presence of c-PAN containing a C_{uc} . The energy change caused by the interaction between the PS and the C_{uc} site as well as the following fragmentation process is around -2.8 eV and therefore more favorable than the previous case, in which the PS fragments break into two S_3 units (-1.2 eV vs -2.8 eV). This behavior is different from another model where no LiTFSI was present in the system and where the S_6 chain was chemically bound to the polymer. In other words, the capturing mechanism can occur in different ways. In the presence of LiTFSI, the salt removes Li from the PS, leading to two solvated Li atoms in the salt. In the abovementioned cases, Li that initially originated from the PS is completely removed by the polymer. However, in the presence of LiTFSI, the Li atom that is solvated by the polymer and by PS is not completely removed from the PS during the simulation time. In the simulations where the salt is not present in the electrolyte, Li absorbs into the polymer backbone and, after a certain period of time, this Li does not coordinate with the PS anymore. Additionally, the final fragments are maintained separated from each other, indicating that the salt can affect the interactions between PS and the polymer. In fact, this indicates that the salt deaccelerates the sulfur chain recombination and the Li capturing by the backbone.

When the electrolyte is composed of DOL and LiTFSI, but LiTFSI is still too far (7.9 Å) from the polymer to interact with the C_{uc} site, the solvent molecules react with the undercoordinated carbon C_{uc} (see Figure 3). In the first step, an O atom from a DOL molecule binds to the C_{uc} at the backbone (~ 0.8 eV); the overcoordination of O atoms from DOL leads to C–O bond cleavage that finally produces $OCH_2CH_2OCH_2$ with O being connected to the C atom from the C_{uc} of the polymer. The total energy change in the decomposition mechanism of DOL is around 2.3 eV. The simulations suggest that PS will be captured in preference to the solvent decomposition (-4.4 eV vs -2.3 eV). This is an indication that if a C_{uc} is formed at the polymer, the C_{uc} will promote solvent decomposition. This might lead to further reactions that require additional analyses and simulations. Regarding the LiTFSI salt, the simulation indicates that it does not tend to react with the C_{uc} of the polymer, since in our simulations even when the salt is placed close (3.1 Å) to the C_{uc} (see Figure S3),

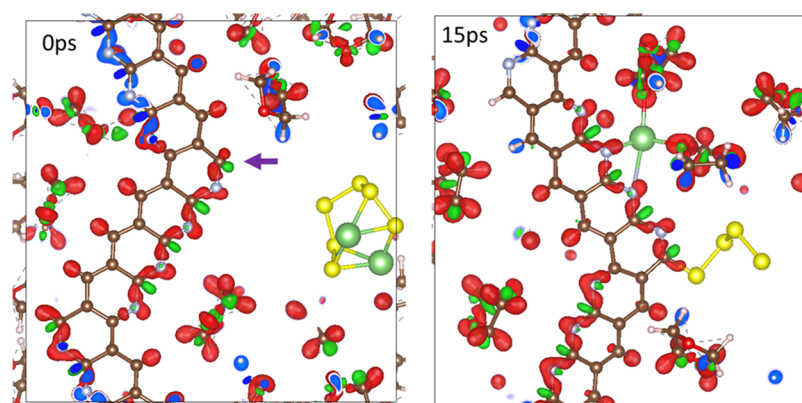


Figure 4. Charge difference isosurface of the interaction between c-PAN and PS. The isosurface has a cutoff of $0.3e$, where red is an accumulation of electrons (negative charge) and green is a depletion of electrons (positive charge). The blue color represents a cleavage in the isosurface produced by the visual plan cleavage. The purple arrow indicates the position of the C_{uc} . For simplicity of the model and image, a slide of the cell is visualized with the atoms and isosurface, while the other regions of the cell are hidden. Color coding: C is brown, N is cyan, H is white, O is red, and S is yellow.

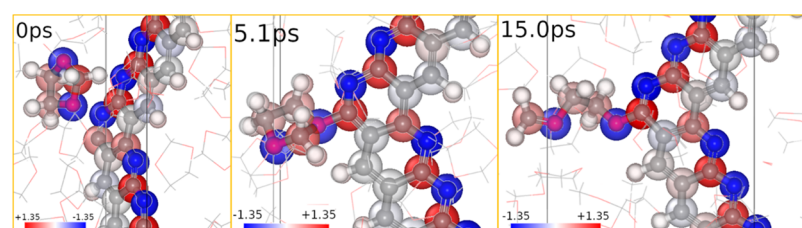


Figure 5. Atom charge distribution of the system as described in Figure 3. The color code is described in Figure 1; however, the atoms are displayed in a translucent way overlapping the charged spheres. The charged sphere is colored accordingly with the scale in this figure.

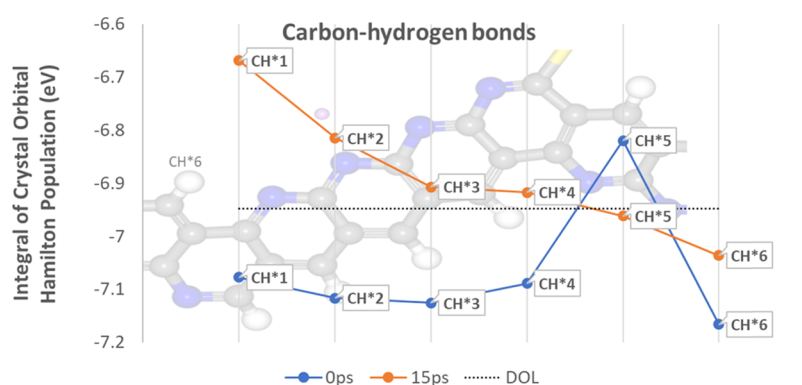


Figure 6. Integral of the Crystal Orbital Hamilton Population of C–H bonds that belong to the system represented in Figure 1A. CH*# is numbered according to the position of each C–H bond, where # is an integer number.

the salt molecule moves away from the polymer backbone, assuming a configuration similar to the one shown in Figure 1B. The solvent indicates to play an important role by removing the available C_{uc} , mainly during the first cycles before the dissolution of PSs. Therefore, changing the solvent has the potential to improve cyclability. However, it cannot be considered conclusive that DOL will always react with C_{uc} because c-PAN can capture Li ions from the PS and change its charge. The decomposition mechanism of DOL in a charged c-PAN should be further investigated to elucidate this mechanism.

To illuminate the interactions and processes of capturing a PS by the c-PAN, we investigated changes in the charge density distribution around the C_{uc} site of c-PAN and after capturing the PS (see Figure 4). In the c-PAN backbone, H

atoms are surrounded by a negative charge (accumulation of electrons); however, H itself tends to have a slightly positive average total charge of $\sim 0.10e$. The N atoms have an average charge of $-1.20e$, while carbon atoms bound to N atoms tend to have a charge of $1.05e$. The negative charge also concentrates on the lone-pair orbital of the N atoms and between C–N bonds, while the positive charge (electron depletion) accumulates on the lone-pair region of C atoms. Moreover, Li prefers staying between two nitrogen lone pairs. The charge of the C atom that forms the C_{uc} site changes slightly from 0.55 to $0.65e$ after binding the PS. During the first 15 ps of the simulation, the total charge of the c-PAN changes from being neutral to $-0.92e$, while the PS changes from -1.61 to $-0.71e$. Consequently, our simulations indicated that the PS transfers electrons to the c-PAN when bound to an

undercoordinated carbon (i.e., C_{uc}) and that the polymer might also act as an active material in a battery by accumulating Li ions. Moreover, in the PS, the charge decays from the beginning of the chain, where the S atom is attached to the polymer, toward the S atom that holds the Li ion, finally becoming more negative on the edge of the chain. The isosurface charges indicate that the PS and Li ions were captured by the lone pair of C_{uc} and N atoms, respectively.

When the electrolyte decomposes prior to binding to the C_{uc} site, there is a more pronounced change in the charge at the undercoordinated carbon as well as the carbon of the DOL ($-CH_2-$) that breaks its bond with O (as shown in Figure 5). The C_{uc} changes its charge from 0.45 to 1.30e, while the latter carbon atom from the electrolyte ($-CH_2-$) reduces its charge from 0.86 to 0.43e. Different from the previous case, in both structures, the decomposed DOL and the polymer remain neutral. Thus, although the polymer C_{uc} leads to solvent decomposition, there is no net charge transfer occurring between the reactants as observed with the capturing of a PS.

The integral of the COHP for C–H bonds that belong to c-PAN is shown in Figure 6. It indicates an increase of the COHP energy after adsorption of the PS by the C_{uc} . In general, the more negative the COHP energies are, the stronger is the bond between the two atoms. Consequently, once the PS is captured by the C_{uc} and Li is removed from the PS to the backbone, the C–H bonds become weaker, changing from around -7.07 eV to around -6.88 eV. The weakest bonds in the c-PAN tend to be localized close to where the Li ion was adsorbed on the polymer (CH*1, CH*2, and CH*3 in Figure 6). This may be associated with the excess of electrons in the polymer backbone after absorption of the PS. The average integral of the COHP for C–H that belongs to the DOL has an average value of -6.94 eV, which tends to be lower than the C–H bonds from c-PAN after having captured the PS (~ -6.88 eV). In other words, DOL tends to be more stable and c-PAN less stable against the removal of hydrogen after adsorption of PS. Therefore, in a competition to remove H atoms from the solvent or the polymer, the absorption of PS and Li in the PAN may affect the species that should react.

As for the cathode materials, where H atoms from the solvent can be removed by salt species,^{47–50} the same may happen to c-PAN during cycling, creating new C_{uc} units. Nevertheless, some DOL molecules have lower COHP energies than the C–H bonds of c-PAN. The smallest COHP for DOL and c-PAN were 6.43 and 6.67 eV, respectively. This may lead to the decomposition of the solvent, instead of the creation of new C_{uc} sites. Moreover, since the reactions may involve depletion of electrons, it is necessary to investigate whether the c-PAN becomes more or less stable during such an electron depletion process.

The previously decomposed solvent molecules (e.g. as shown in Figure 3) might also trigger removal of H from c-PAN or supply H for the sulfur atoms in the SPAN. Further analyses should be done to better understand the behavior of c-PAN under such conditions. Here, our calculations indicate that the stability of the C–H bonds in c-PAN should change during cycling of the LiSB. Assuming that the minimum COHP (C–H) value is more relevant than the overall average value, other solvents with stronger C–H bonds could be used instead of DOL to activate the generation of C_{uc} during the LiSB cycling instead of decomposing the solvent (when salts such as LiBr are used in the electrolyte).

We also calculated the bond strength of PS as a fragment (observed after 0.76 ps) and after recombination of the S–S bond (observed after 15 ps), as shown in Figure 2. The results summarized in Figure S4 suggest that in the fragmented structure, the bond strength of both S–Li and S–S bonds is higher compared to that of the recombined case. The COHP of the sulfur fragment varies from an average of -6.90 to -5.56 eV, and the COHP of every S–Li bond changes from -0.77 to -0.27 eV, on comparing simulation snapshots at 0.76 and 15.0 ps, respectively. Nevertheless, S–Li bonds are typically weaker than S–S bonds (~ 5.7 eV vs 0.6 eV); the covalent and ionic effect on COHP can also be observed in the literature for Na–S, Li–O, and S–S.^{97–99} Therefore, it is expected that fragments might recombine as a whole, without encountering further cleavage of the sulfur chains, which agrees with our simulations.

CONCLUSIONS

In the present work, we have studied the capturing process of a PS by a C_{uc} in the c-PAN backbone, the reactions between this C_{uc} and the solvent, and the effect of the salt in the PS-capturing process. We considered undercoordinated carbon (C_{uc}) that was formed by three different processes:

- during the syntheses, where H is removed from the backbone by sulfur, without binding a sulfur chain to C_{uc} ;
- by removal of sulfur from carbon during battery cycling; and
- by removal of hydrogen from the backbone due to its reaction with a nearby salt molecule.

Our simulations indicate that c-PAN is capable of catching a PS chain dissolved in the electrolyte and of removing Li from this PS when C_{uc} is available at the polymer backbone. In other words, when a carbon atom at the c-PAN structure is missing its H termination, this carbon can bind to the sulfur-chain of a PS. During this capturing process, some S–S bonds might break, forming temporary fragments attached to the adsorbed PS. Afterward, these PS might recombine, forming S–S bonds, while the PS will assume a linear configuration with only one Li-ion being attached to the edge of the PS, and the other Li ion being adsorbed at N sites of the polymer. Consequently, c-PAN might also act as an active material in the battery. The PS takes a similar configuration when the PS is initially attached to the polymer or when being solvated in the electrolyte. However, the salt seems to influence the lifetime of the PS fragments, increasing the time required for fragment recombination and leading to the formation of longer chains. Moreover, the PS shows a tendency to weaken the bond between hydrogen and the backbone. The mechanism in which the C_{uc} captures the PSs may change accordingly with the initial configuration of the PS. However, the recombination of S–S chains suggests that the attached PS will commonly assume a linear configuration with one Li ion on the edge of the sulfur chain.

These C_{uc} sites are also able to react with the solvent, breaking a C–O bond (from carbon on $-CH_2-$) and creating a new branch in the c-PAN polymeric structure. Further research may investigate if PS can absorb on the decomposed DOL ramification, bounded by the c-PAN or further reactions can occur in that ramification. Also, more investigation should be conducted to investigate whether salt species (such as LiBr) are capable to generate additional C_{uc} units. The absorption of

PS leads to weaker C–H bonds in the c-PAN, when closer to the attached PS. However, c-PAN has on average weaker C–H bonds than DOL, while the solvent has the weakest C–H bonds.

We can conclude that c-PAN might also act as an active material, not only by allowing lithiation of sulfur but also by capturing Li ions. The polymer backbone can capture PS from the electrolyte, producing PS fragments. These PS fragments might recombine to form long-chain PSs. However, in the presence of the salt, this recombination process is decelerated. Finally, the strength of C–H bonds decreases with the degree of lithiation of the polymer.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acsomega.0c06240>.

Complementary AIMD figures mentioned above and the COHP analyses of S–S and S–Li (PDF)

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Notes

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